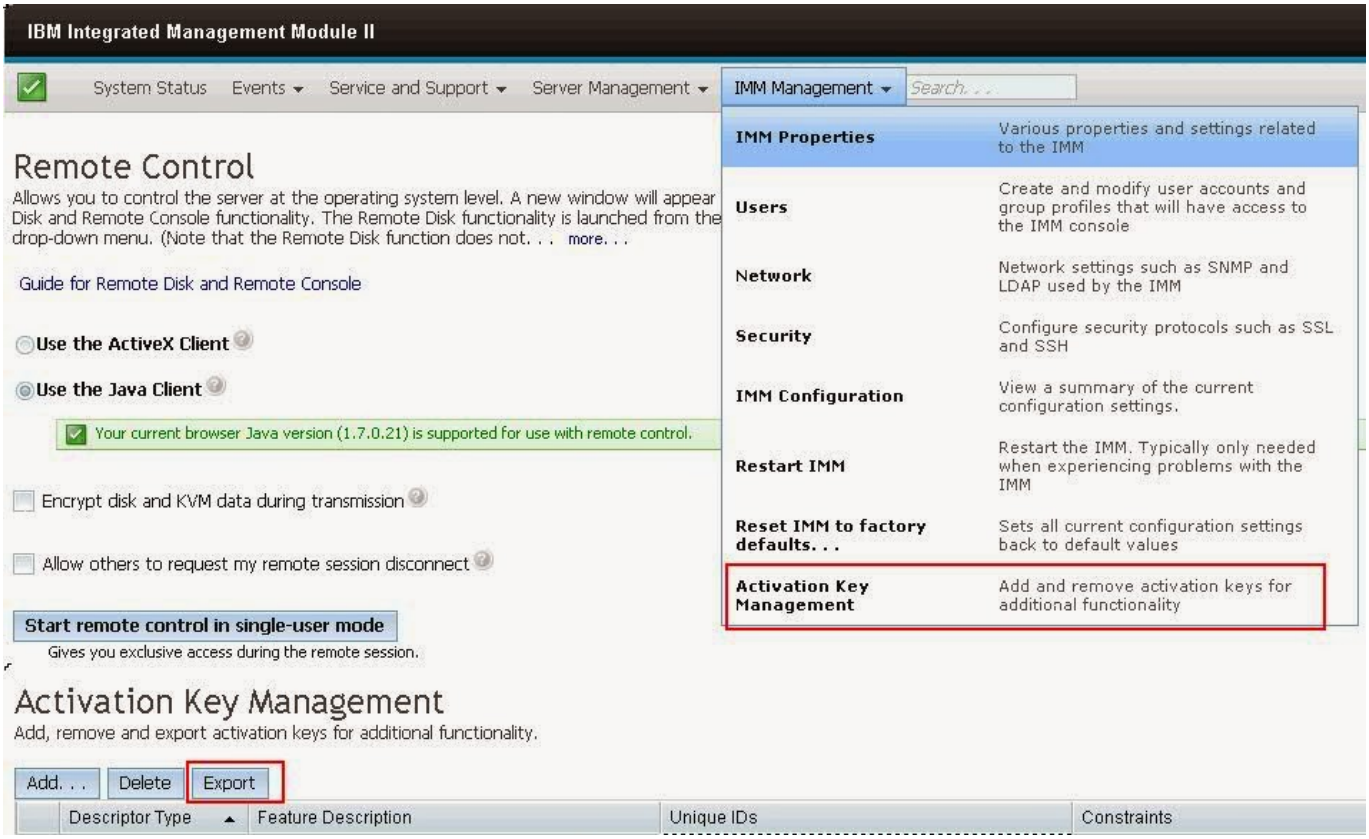
Ibm Imm Activation Key 124



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activation energy, Activation Energy, Activation of energy, Activation Energy. Energy is a substance which is required to cause a chemical reaction to take place. The activation energy (E) is the energy required to cause a chemical reaction to take place. If this energy is low, the reaction will take place readily; if the energy is high, the reaction will be difficult to take place. The activation energy, E, is the quantity of energy required to cause the chemical reaction to take place. Some chemical reactions can be as simple as two atoms combining to form a different atom (e.g. H2O to H2 + O2). Other reactions require many components (e.g. an enzyme helping a molecule break down, or a catalyst enabling the reaction to take place at a lower temperature or in a specific environment). An example of a simple reaction is the combustion of a hydrocarbon. A combustion reaction consists of two reactants: the hydrocarbon (carbon), and air (oxygen). The only difference between these two is that a fuel contains a fuel which is the energy source, and an oxidizer contains oxygen which is needed to burn the fuel to produce energy. When the mixture of fuel and oxidizer is ignited, they release large amounts of energy in the form of light, heat and other forms of energy (i.e. work). Chemical reactions are associated with a change in the structure of the reactants, or their products. For example, the combustion reaction has a net energy gain of 62 kJ/mol. The reaction involves the breaking of a bond between carbon and oxygen, and the formation of two other bonds. This change in the structure of the reactants is called bond activation, and the energy required for this change is called the activation energy. An enzyme is an example of an activator. It speeds up the reaction by lowering the activation energy, but is not needed for the reaction to take place. The activation energy is sometimes called the reactant's activation energy. For some reactions, the same quantity of energy is needed to trigger the reaction as there is to drive the reaction forward. In such cases, the activation energy is known as the transition state energy or activation energy. In other cases, the same quantity of energy is needed to trigger the reaction as there is to drive the reaction forward. In such cases, the activation energy is known as the transition state energy or activation energy. For reactions requiring thermal activation energy, the reaction requires a specific temperature to occur (temperature dependency). This activation 82157476af

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